**Ksp Homework**

**Problem #1:** Determine the Ksp of silver bromide, given that its molar solubility is 5.71 x 10¯7  moles per liter.

**Problem #2:** Determine the Ksp of calcium fluoride (CaF2), given that its molar solubility is 2.14 x 10¯4  moles per liter.

**Problem #3:** Determine the Ksp of mercury(I) bromide (Hg2Br2), given that its molar solubility is 2.52 x 10¯8 moles per liter.

**Problem #4:** Calculate the Ksp for Ce(IO3)4, given that its molar solubility is 1.80 x 10¯4 mol/L

**Problem #5:** Calculate the Ksp for Mg3(PO4)2, given that its molar solubility is 3.57 x 10-6 mol/L.

**Problem #6:**  Compound = AgCN, Solubility = 7.73 x 10¯9 M. Calculate Ksp.

**Problem #7**: The Ksp of AgBr is 3.2 x10-13. Calculate the molar solubility.

**Problem #8**: The Ksp of PbI2 is 8.7 x10-9 Calculate the solubility.

**Problem #9**: . Calculate the solubility of Zn(OH)2 in mg/L. Ksp = 4.5 x10-17

(hint: 1 gram = 1000 mg)