**Lewis Dot, Formula Unit & Naming Practice Sheet**

**Notes:**

1. An **ionic bond** is an attraction of a *cation* for an *anion* resulting from the transfer of electrons. Remember, the smaller nonmetals are more electronegative and pull the electrons close, away from the larger, less electronegative metals.
2. When naming ionic compounds, the Metal is named first, followed by the nonmetal with an –ide ending. *Ex. Sodium Fluorine becomes Sodium Fluoride.*
3. **Formula Unit:** Lowest whole number ratio of elements in the compound. Ex. Ca3N2

|  |  |
| --- | --- |
| 1. Draw the Lewis Structure for Mg & ClFormula Unit: \_\_\_\_\_\_\_\_\_Name of Compound:  | 2. Draw the Lewis Structure for Mg & SFormula Unit: \_\_\_\_\_\_\_\_\_Name of Compound:  |
| 3. Draw the Lewis Structure for K & FFormula Unit: \_\_\_\_\_\_\_\_\_Name of Compound: | 4. Draw the Lewis Structure for K & OFormula Unit: \_\_\_\_\_\_\_\_\_Name of Compound: |
| 5. Draw the Lewis Structure for Be & NFormula Unit: \_\_\_\_\_\_\_\_\_Name of Compound: | 6. Draw the Lewis Structure for Ca & PFormula Unit: \_\_\_\_\_\_\_\_\_Name of Compound: |
| 7. Draw the Lewis Structure for Al & FFormula Unit: \_\_\_\_\_\_\_\_\_Name of Compound: | 8. Draw the Lewis Structure for Ca & IFormula Unit: \_\_\_\_\_\_\_\_\_Name of Compound: |
| 9. Draw the Lewis Structure for Rb & OFormula Unit: \_\_\_\_\_\_\_\_\_Name of Compound: | 10. Draw the Lewis Structure for Sr & FFormula Unit: \_\_\_\_\_\_\_\_\_Name of Compound: |
| 11. Draw the Lewis Structure for Al & ClFormula Unit: \_\_\_\_\_\_\_\_\_Name of Compound: | 12. Draw the Lewis Structure for Mg & PFormula Unit: \_\_\_\_\_\_\_\_\_Name of Compound: |
| 13. Draw the Lewis Structure for B & OFormula Unit: \_\_\_\_\_\_\_\_\_Name of Compound: | 14. Draw the Lewis Structure for Be & SFormula Unit: \_\_\_\_\_\_\_\_\_Name of Compound: |